Counce, R. M., and J. J. Perona, "Gaseous Nitrogen Oxide Absorption in a Sieve Plate Column," Ind. Eng. Chem. Fundam., 18, 400 (1979). Denbigh, K. G., and A. J. Prince, "Kinetics of Nitrous Gas Absorption in Aqueous Nitric Acid," J. Chem. Soc., 53, 790 (1947).

Greig, J. A., and P. G. Hall, "Thermal Oxidation of Nitric Oxide at Low Concentrations," Trans. Faraday Soc., 63, 655 (1967).

Kameoka, Y., and R. L. Pigford, "Absorption of Nitrogen Dioxide into Water, Sulfuric Acid, Sodium Hydroxide, and Alkaline Sodium Sulfite Aqueous Solution," Ind. Eng. Chem. Fundam., 16, 163 (1977).

Kobayashi, H., N. Takezawa and T. Niki, "Removal of Nitrogen Oxides with Aqueous Solutions of Inorganic and Organic Reagents," Environ. Sci. Tech., 11, 190 (1977). Levenspiel, O., "Chemical Reaction Engineering," 2nd ed., p. 72, John

Wiley, New York (1972).

Peters, M. S., C. P. Ross, and J. E. Klein, "Controlling Mechanism in the Aqueous Absorption of Nitrogen Oxides," AIChE J., 1, 105

Sherwood, T. K., "Absorption and Extraction," McGraw Hill, New York (1937).

Treharne, R. W., M. R. Bruce, D. R. Moles, and C. K. McKibben, Technical Note I, Charles F. Kettering Research Laboratory, 1978.

Waesser, B., "The Atmosphere Nitrogen Industry," Vols. 1 and 2, P.

Blakiston's and Sons, Philadelphia, Pennsylvania, 1926. Wendel, M. M., and R. L. Pigford, "Kinetics of Nitrogen Tetroxide Absorption in Water," AIChE J., 4, 249 (1958).

Manuscript received January 23, 1980; revision received April 23, and accepted May 16,

Calculation of Critical Points from Cubic Two-Constant **Equations of State**

MICHAEL L. MICHELSEN

Instituttet for Kemiteknik The Technical University of Denmark Lundtoftevei 100, Building 229 DK 2800 Lyngby, Denmark

and

ROBERT A. HEIDEMANN

Department of Chemical Engineering The University of Calgary 2920-24 Ave. N.W. Calgary, Alberta T2N-1N4, Canada

Evaluation of critical points for multicomponent mixtures based on an equation of state has attracted considerable attention in recent years. The first general procedure for direct determination of critical temperature and pressure was developed by Peng and Robinson (1977), who used a criterion based on the Gibbs free energy, while a later method by Heidemann and Khalil (1979) is based on the Helmholtz free energy. Indirect methods, where the critical point is found from a phase envelope construction, have also been suggested (Assilineau et al, 1979, Michelsen, 1980).

The direct methods result in two complex non-linear equations. The method of Heidemann and Khalil is far superior, since it requires a computational effort proportional to n^3 (n being the number of components) as compared to n^5 in the procedure of Peng and Robinson. The partial derivatives required using the Helmholtz free energy are much more rapidly evaluated than those using the Gibbs free energy.

The present paper describes a computational modification of the method of Heidemann and Khalil, applicable to simple, two-constant Cubic Equations of State such as the SRKequation (Soave, 1972) and the PR-equation (Peng and Robinson, 1976). This modification reduces the computing time by a factor of 5-10.

For systems where the unlike binary interaction parameters may be calculated from the pure-component parameters using the geometric mean values, a further simplification is given, whereby computing times are reduced to a few milliseconds, almost irrespective of the number of components.

Method of Heidemann and Khalil

Using a Taylor series expansion of the Helmholtz free energy, Heidemann and Khalil derived the following criterion: At the critical point a mixture of total composition N must satisfy

$$\mathbf{Q} \mathbf{\Delta} \mathbf{N} = \mathbf{0}, \ \mathbf{\Delta} \mathbf{N}^T \mathbf{\Delta} \mathbf{N} = 1 \tag{1}$$

0001-1541-81-4388-0519-\$2.00. CThe American Institute of Chemical Engineers,

and

$$C = \sum_{i} \sum_{j} \sum_{K} \Delta N_{i} \Delta N_{j} \Delta N_{K} \left(\frac{\partial^{3} A}{\partial N_{j} \partial N_{j} \partial N_{K}} \right)_{T,V} = 0$$
 (2)

where A is the Helmholtz' free energy, and

$$Q_{ij} = \left(\frac{\partial^2 A}{\partial N_i \partial N_i}\right)_{T,V} = RT \left(\frac{\partial \ln f_i}{\partial N_i}\right)_{T,V} \tag{3}$$

To evaluate T_c and v_c nested iterations were used. Newton iteration at a fixed value of v is used to determine a temperature where the set of homogeneous equations (1) have a nontrivial solution. The elements ΔN_i are calculated, and evaluation of the cubic form C is used to correct v in an outer loop. For simple cubic equations of state the evaluation of C requires essentially only one or two double summations, and for general equations numerical evaluation as described in (Michelsen, 1980) is inexpensive. The computationally expensive part is that of solving (1). The solution technique suggested in the following takes advantage of the specific temperature dependence in the SRKand PR-equations of state.

Two-Constant Cubic Equations of State

We shall consider Cubic Equations of State of the form

$$P = \frac{RT}{v - b} - \frac{a}{(v + \delta_1 b)(v + \delta_2 b)} \tag{4}$$

where the mixture constants a and b are given by the mixing rules

$$a = \sum_{i} \sum_{j} y_i y_j a_{ij} \tag{5}$$

$$b = \sum_{i} y_i b_i \tag{6}$$

$$a_{ij} = (a_{ii}a_{jj})^{\frac{1}{2}}(1 - k_{ij}) \tag{7}$$

The pure component constants are found from

$$a_{ii} = \Omega_0 R^2 T c_i^2 / P c_i [1 + m_i (1 - (T/Tc_i)^{\frac{1}{2}})]^2$$
 (8)

 $m_i = m(\omega_i)$

$$b_i = \Omega_b R T c_i / P c_i \tag{9}$$

The constants Ω_a and Ω_b are determined applying eq. (4) at the critical point for a pure component.

For the SRK-equation of state

$$\delta_1 = 1, \ \delta_2 = 0, \ \Omega_a = 0.42748, \ \Omega_b = 0.08664$$
 (10)

$$m(\omega_i) = 0.480 + 1.574\omega_i - 0.176\omega_i^2 \tag{11}$$

and for the Peng-Robinson equation

$$\delta_1 = 1 + \sqrt{2}, \, \delta_2 = 1 - \sqrt{2},$$

$$\Omega_a = 0.45724, \ \Omega_b = 0.07780 \ (12)$$

and

$$m(\omega_i) = 0.37464 + 1.54226\omega_i - 0.26992\omega_i^2$$
 (13)

The elements Q_{ij} are evaluated from

$$n_{T}Q_{ij} = n_{T}RT\left(\frac{\partial \ln f_{i}}{\partial N_{j}}\right)_{T,V}$$

$$= RT\left\{\frac{\delta_{ij}}{y_{i}} + (\beta_{i} + \beta_{j})F_{1} + \beta_{i}\beta_{j}F_{1}^{2}\right\}$$

$$+ \frac{a}{h}\left\{\beta_{i}\beta_{j}F_{3} - a_{ij}/a F_{5} + (\beta_{i}\beta_{j} - \alpha_{i}\beta_{j} - \alpha_{j}\beta_{i})F_{6}\right\}$$
(14)

where

$$\beta_i = \frac{b_i}{b}, \ \alpha_i = \frac{1}{a} \sum_i y_i a_{ij} \tag{15}$$

and where $F_1 - F_8$ given in Appendix A are simple functions of the dimensionless volume $\kappa = v/b$.

From the temperature dependence of a_{ii} it is evident that all elements of Q_{ij} are second degree polynomials in $\lambda = T^{\frac{1}{2}}$, that is

$$\mathbf{O} = \mathbf{O}_{o}(\kappa) + \lambda \mathbf{O}_{1}(\kappa) + \lambda^{2} \mathbf{O}_{2}(\kappa)$$
 (16)

Hence, to solve for T and ΔN at a given value of κ is equivalent to solving the generalized algebraic eigenvalue problem

$$(\mathbf{O}_{a} + \lambda \mathbf{O}_{1} + \lambda^{2} \mathbf{O}_{2}) \Delta \mathbf{N} = 0 \tag{17}$$

Since we are interested in only one eigenvalue/eigenvector pair (that corresponding to the largest temperature root) the inverse iteration technique described in (Wilkinson, 1965, Chap. IX, sect. 61) is very attractive. A trial value of λ is chosen, and triangular decomposition of $\mathbf{Q}(\kappa,\lambda)$ is performed. Each subsequent iteration requires only inexpensive matrix-vector multiplications. The convergence is linear, with a rate inversely proportional to the error in our trial value of λ .

Treating equation (1) as the eigenvalue problem of equation (17) is the principal departure from the procedure of Heidemann and Khalil (1979) which we propose.

The cubic form C can for a given ΔN be evaluated from

$$n_T^2 C = RT \left\{ -\sum_i \Delta N_i^3 / y_i^2 + 3\overline{N} (\overline{\beta} F_1)^2 + 2(\overline{\beta} F_1)^3 \right\} + \frac{a}{b} \left\{ 3\overline{\beta}^2 (2\overline{\alpha} - \overline{\beta}) (F_3 + F_6) - 2\overline{\beta}^3 F_4 - 3\overline{\beta} \overline{\alpha} F_6 \right\}$$
(18)

where

$$\overline{N} = \sum_{i} \Delta N_{i}, \ \overline{\beta} = \sum_{i} \Delta N_{i} \beta_{i}$$
 (19)

$$\overline{\alpha} = \sum_{i} \Delta N_{i} \alpha_{i}$$
, and $\overline{a} = \frac{1}{a} \sum_{j} \sum_{i} \Delta N_{i} \Delta N_{j} a_{ij}$ (20)

The simplicity of evaluating (18) once λ and ΔN have been found, will be obvious.

At a critical point, in addition to equation (17), we must also have $n_T^2C=0$. A new procedure is proposed here for satisfying these equations.

Let the eigenvalue and eigenvector corresponding to our initial choice of κ , $\kappa = \kappa_1$ be $(\lambda_1, \Delta N_1)$. The vector is normalized, $(\Delta N_1)^T (\Delta N_1) = 1$, and the cubic form $C_1 = C(\lambda_1, \Delta N_1)$ is evaluated.

An increment in κ is chosen, $\kappa_2 = \kappa_1 + \Delta \kappa$, and as initial estimates for λ_2 and ΔN_2 are chosen

$$\lambda_2 = \lambda_1 + \left(\frac{d\lambda}{d\kappa}\right) \Delta \kappa, \ \Delta N_2 = \Delta N_1$$
 (21)

where $\left(\frac{d\lambda}{d\kappa}\right)$ is evaluated from

$$\frac{d\lambda}{d\kappa} = -(\Delta \mathbf{N}^T \mathbf{Q}_{\kappa} \Delta \mathbf{N}) / (\Delta \mathbf{N}^T \mathbf{Q}_{\lambda} \Delta \mathbf{N})$$
 (22)

with

$$\mathbf{Q}_{\kappa} = \frac{\partial}{\partial \kappa}(\mathbf{Q})$$
 and $\mathbf{Q}_{\lambda} = \frac{\partial}{\partial \lambda}(\mathbf{Q})$

Evaluation of C_2 then permits adjustment of κ for subsequent iterations by inverse interpolation. Following the recommendations of Heidemann and Khalil, C is replaced by $C^* = C(\kappa - 1)^2$, and the iteration scheme becomes

$$\kappa_{k+1} = \kappa_k + \frac{C_k^*}{C_{k-1}^* - C_k^*} (\kappa_k - \kappa_{k-1})$$
(23)

To avoid discrete jumps in ΔN the additional requirement

$$(\mathbf{\Delta}\mathbf{N}_k)^T \mathbf{\Delta}\mathbf{N}_1 > 0 \tag{24}$$

is imposed in the normalization of ΔN_k . As initial estimates are used

$$\kappa_1 = 3.5; \ \lambda_1^2 = 1.3 \sum_i y_i T c_i; \ \Delta N_1 = N$$
(25)

Normally 5-7 iterations in κ are required. At κ_1 about 4-6 iterations are necessary to determine λ_1 , while all subsequent inner iterations usually converge in 1 or 2 steps, since an excellent initial estimate of λ and ΔN is available.

The total computing time for the 32 systems investigated by Peng and Robinson (1977) is 0.4 sec. (IBM 3033, FORTG compiler), ranging from 0.006 sec. for 3 components to about 0.07 sec. for 12 components.

Systems with Zero Interaction Coefficients

The method described above can be directly extended to an equation of state with two interaction deviation parameters, as described in (Heidemann and Khalil, 1979). Here we shall describe an alternative development which can be used for systems where all k_{ij} are equal to zero. Such systems are of importance since the assumption $k_{ij} = 0$ is frequently used for mixtures of pure hydrocarbons.

Provided all k_{ij} are equal to zero, we obtain

$$a_{ij} = (a_{ii}a_{jj})^{\frac{1}{2}} = \alpha_i \alpha_j a \tag{26}$$

Equation (1) can be written

$$S_i = \sum_i Q_{ij} \Delta N_j = 0, \ i = 1, 2, \dots, n$$
 (27)

or, using (13), (14), (18), and (19)

$$RT\left(\frac{\Delta N_{i}}{y_{i}} + (\beta_{i}\overline{N} + \overline{\beta})F_{1} + \beta_{i}\overline{\beta}F_{2}\right) + \frac{a}{b}\left\{\beta_{i}\overline{\beta_{i}}F_{3} + (\beta_{i}\overline{\beta} - \alpha_{i}\overline{\beta} - \beta_{i}\overline{\beta})F_{6} - \alpha_{i}\overline{\alpha}F_{5}\right\} = 0 \quad (28)$$

From (27) the following set of conditions must also hold

$$\sum_{i} y_i S_i = 0; \quad \sum y_i \beta_i S_i = 0; \quad \sum y_i \alpha_i S_i = 0$$
 (29)

Substitution of (27) results in a homogeneous set of only 3 equations in $(\overline{N}, \overline{\alpha}, \overline{\beta})$. Elimination of \overline{N} yields after some rearrangement the following form

$$\binom{D_1 \ D_2}{D_2 \ D_3} \left(\frac{\overline{\beta}}{\alpha}\right) = \binom{0}{0} \tag{30}$$

where

$$S_{\beta\beta} = \sum_{i} y_{i} \beta_{i}^{2}, \ S_{\alpha\beta} = \sum_{i} y_{i} \alpha_{i} \beta_{i}, \ S_{\alpha\alpha} = \sum_{i} y_{i} \alpha_{i}^{2}$$
 (31)

$$D_1 = RT(1 + F_1) + \frac{a}{b} \{ S_{\beta\beta} F_8 + (S_{\beta\beta} - S_{\alpha\beta}) F_6 \}$$
 (32)

$$D_2 = \frac{a}{b} \{ S_{\beta\beta} F_7 + (S_{\beta\beta} - S_{\alpha\beta}) F_5 \}$$
 (33)

$$D_3 = RT F_1 + \frac{a}{b} \{ S_{\alpha\beta} F_8 + (S_{\alpha\beta} - S_{\alpha\alpha}) F_6 \}$$
 (34)

The temperature root is found from

$$D_1 D_4 - D_2 D_3 = 0 (36)$$

Since D_1 , D_2 , D_3 , and D_4 are all second degree polynomials, in λ = $T^{\frac{1}{2}}$ we get 4 and only 4 nonzero roots. Solving (36) and using for normalization $\alpha = 1$ yields

$$\overline{\beta} = -D_2/D_1 \tag{37}$$

$$RT\overline{N} + \frac{a}{h} F_7 \overline{\alpha} + (RTF_1 + \frac{a}{h} F_8) \overline{\beta} = 0$$
 (38)

The individual ΔN_i can be recovered from (28), and all terms needed for evaluation of *C* are available, since for $k_{ij} = 0$, $\bar{a} = \bar{\alpha}^2$

Computing times using this extremely simple result are of the order 2-4 millisec, almost independent of the number of components in the mixture.

For systems with $k_{ij} \neq 0$ critical properties calculated using the $k_{ij} = 0$ assumption often deviate only slightly from those found using actual k_{ij} -values and serve as excellent initial estimates for the complete calculation. Hereby total computing times can be reduced approximately by a factor of 2.

Conclusion and Suggestions

The development of rapid and accurate procedures for calculation of critical points is of value for general phase equilibrium calculations, since convergence is often difficult to obtain in the critical region. The ability to locate the critical region in advance will greatly help selecting the proper computational strategy.

The reduction in dimensionality in the case when all binary interaction coefficients are zero can also be utilized in phase equilibrium calculations for multicomponent systems. Using a and b as variables rather than using phase compositions reduces the number of independent variables to three for phase envelope calculations and for flash calculations at specified T and P, irrespective of the number of components. Computational methods based on this simplification are currently being investigated.

APPENDIX: Volume functions $F_1(\kappa) - F_8(\kappa)$, $\kappa = v/b$:

$$F_1 = 1/(\kappa - 1)$$

$$F_2 = 2\{(\delta_1/(\kappa + \delta_1)) - (\delta_2/(\kappa + \delta_2))\}/(\delta_1 - \delta_2)$$

$$F_{3} = \{(\delta_{1}/(\kappa + \delta_{1}))^{2} - (\delta_{2}/(\kappa + \delta_{2}))^{2}\}/(\delta_{1} - \delta_{2})$$

$$F_{4} = \{(\delta_{1}/(\kappa + \delta_{1}))^{3} - (\delta_{2}/(\kappa + \delta_{2}))^{3}\}/(\delta_{1} - \delta_{2})$$

$$F_{5} = 2 \ln\{(\kappa + \delta_{1})/(\kappa + \delta_{2})\}/(\delta_{1} - \delta_{2})$$

$$F_{6} = F_{2} - F_{5}$$

$$F_{7} = -F_{2}/(1 + F_{1})$$

$$F_{8} = F_{3}/(1 + F_{1})$$

NOTATION

= Helmholtz free energy

a, b= Parameters in equation of state = Cubic form, eq. 2 \boldsymbol{C}

 $D_1 - D_4 = \text{See eqs. } 32-34$

Volume functions, see Appendix

= Fugacity of component i

= Binary interaction coefficient

 $\frac{N}{N_i}$ = Component vector

= Number of moles of i'th component

Parameter in equation of state

= Number of components in mixture

Total number of moles in mixture n_T

Pressure

 $\frac{Q}{R}$ Matrix of partial derivatives

Gas constant S_i See eq. 27

TAbsolute temperature

Total volume

Molar volume

Mole fraction, component i

Subscripts

Critical property i, j, k = Component numbers

Greek Letters

= See eq. 15

= See eq. 15

 $\delta_1, \, \delta_2$ = Constants in equation of state

= Dimensionless volume, $\kappa = v/b$

 Ω_a , Ω_b = Constants in equation of state

= Acentric factor

LITERATURE CITED

Asselineau, L., Bogdanic, G., and J. Vidal: A versatile algorithm for calculating vapour-liquid equilibria, Fluid Phase Equilibria 3(4), 273

Heidemann, R. A. and A. M. Khalil: The calculation of critical points, Paper presented at the 86'th National A.I.Ch.E. Meeting, Houston, April, 1979. AIChE J., 26, 769 (1980).

Michelsen, M. L.: Calculation of phase envelopes and critical points for multicomponent mixtures, Fluid Phase Equilibria 4, I (1980).

Peng, D.-Y. and D. B. Robinson: A new two constant equation of state, Ind. Eng. Chem. Fundamentals 15, 59 (1976).

Peng, D.-Y. and D. B. Robinson: A rigorous method for predicting the critical properties of multicomponent systems from an equation of state, A.I.Ch.E. J. 23, 137 (1977).

Soave, G.: Equilibrium constants from a modified Redlich-Kwong equation of state, Chem. Eng. Sci. 27, 1197 (1972).

Wilkinson, J. H.: The Algebraic Eigenvalue Problem, Clarendon Press, Oxford (1965).

Manuscript received March 20, 1980; revision received June 12 and accepted June 19,